Vrite your name here Surname	Other	names
Pearson Edexcel GCSE	Centre Number	Candidate Number
Chemistry/	Additions	I Science
Unit C2: Discoveri		
	ng Chemistry	Higher Tier Paper Reference 5CH2H/01

Instructions

- Use **black** ink or ball-point pen.
- Fill in the boxes at the top of this page with your name, centre number and candidate number.
- Answer all questions.
- Answer the questions in the spaces provided
 - there may be more space than you need.

Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets
 - use this as a guide as to how much time to spend on each question.
- Questions labelled with an **asterisk** (*) are ones where the quality of your written communication will be assessed
 - you should take particular care with your spelling, punctuation and grammar, as well as the clarity of expression, on these questions.

Advice

- Read each question carefully before you start to answer it.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ▶

PEARSON

The Periodic Table of the Elements

0 4 	20 Ne neon	40 Ar argon 18	84 Kr krypton 36	131 Xe xenon 54	[222] Rn radon 86	t fully
7	19 F fluorine 9	35.5 CI chlorine 17	80 Br bromine 35	127 	[210] At astatine 85	orted but no
9	16 O oxygen 8	32 S sulfur 16	79 Se selenium 34	128 Te tellurium 52	[209] Po polonium 84	ive been rep
5	14 N nitrogen 7	31 P phosphorus 15	75 As arsenic 33	122 Sb antimony 51	209 Bi bismuth 83	rs 112-116 ha authenticated
4	12 C carbon 6	28 Si silicon 14	73 Ge germanium 32	119 Sn tin 50	207 Pb lead 82	omic number:
ဧ	11 B boron 5	27 AI aluminium 13	70 Ga gallium 31	115 In indium 49	204 TI thallium 81	Elements with atomic numbers 112-116 have been reported but not fully authenticated
·			65 Zn zinc 30	112 Cd cadmium 48	201 Hg mercury 80	Elem
			63.5 Cu copper 29	108 Ag silver 47	197 Au gold 79	Rg roentgenium 111
			59 Ni nickel 28	106 Pd palladium 46	195 Pt platinum 78	[271] Ds damstactium 110
			59 Co cobatt 27	103 Rh rhodium 45	192 Ir iridium 77	[268] Mt meitnerium 109
1 Hydrogen			56 Fe iron 26	101 Ru ruthenium 44	190 Os osmium 76	[277] HS hassium 108
			55 Mn manganese 25	[98] Tc technetium 43	186 Re rhenium 75	[264] Bh bohnium 107
	mass ɔol ıumber		52 Cr chromium 24	96 Mo molybdenum 42	184 W tungsten 74	[266] Sg seaborgium 106
Key	relative atomic mass atomic symbol _{rame} atomic (proton) number		51 V vanadium 23	93 Nb niobium 41	181 Ta tantalum 73	[262] Db dubnium 105
	relati atc atomic		48 Ti titanium 22	91 Zr zirconium 40	178 Hf hafnium 72	[261] Rf rutherfordium 104
			45 Sc scandium 21	89 Y yttrium 39	139 La* lanthanum 57	[227] Ac* actinium 89
2	9 Be beryllium 4	24 Mg magnesium 12	40 Ca calcium 20	88 Sr strontium 38	137 Ba barium 56	[226] Ra radium 88
_	7 Li lithium 3	23 Na sodium 11	39 K potassium	85 Rb rubidium 37	133 Cs caesium 55	[223] Fr francium 87

^{*} The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.

Questions begin on next page.

Answer ALL questions

Some questions must be answered with a cross in a box \boxtimes . If you change your mind about an answer, put a line through the box \boxtimes and then mark your new answer with a cross \boxtimes .

The periodic table

1 The positions of ten elements are shown in the outline periodic table.

The letters are not the symbols of the elements.

1 2

3 4 5 6 7 0

E				-			G	J		
L							М			
					Q				R	
Т							X			Z

- (a) Use the letters given in the table to identify
 - (i) the element that is a noble gas.

(1)

(ii) three elements that are in the same period.

(1)

(b) Which element reacts with hydrogen to form a compound that dissolves in water to give an acidic solution?

Put a cross (\boxtimes) in the box to show your answer.

(1)

- A element G
- B element J
- C element R
- D element Z



(c) Which elements, when combined, form an ionic compound?	
Put a cross (☒) in the box to show your answer.	(1)
■ A E and R	(1)
■ B G and J	
□ C G and R	
☑ D R and Z	
(d) Element Q is a metal and a good conductor of electricity.	
Explain how metals conduct electricity.	
	(2)
(e) Reactivity increases down group 1 of the periodic table.	
Elements E and T are in group 1 of the periodic table.	
Explain why element T is more reactive than element E .	
h	(2)
(Total for Question 1	= 8 marks)



Sodium sulfate

- **2** Sodium sulfate is an ionic solid which is soluble in water.
 - (a) Which row of the table shows the likely properties of sodium sulfate?

Put a cross (\boxtimes) in the box to show your answer.

(1)

	melting point / °C	conducts electricity when dissolved in water
⊠ A	22	yes
⊠ B	22	no
⊠ C	884	yes
⊠ D	884	no

- (b) Sodium sulfate contains sodium ions, $\mathrm{Na^{+}}$, and sulfate ions, $\mathrm{SO_4^{\;2^{-}}}$.
 - (i) Complete the sentence by putting a cross (☒) in the box next to your answer.

(1)

The formula for sodium sulfate is

- A NaSO₄
- \square **B** Na(SO₄)₂
- ☑ C Na₂SO₄
- \square **D** Na₂(SO₄)₂
- (ii) Sodium ions can be identified using a flame test.

Which colour in the flame shows the presence of sodium ions?

Put a cross (☒) in the box to show your answer.

(1)

- 🛛 A lilac
- **B** red
- **D** yellow



C) When sodium sulfate solution reacts with barium hydroxide solution, a white precipitate of barium sulfate forms in the mixture. The barium sulphate precipitate can be separated from the mixture and purified. (i) Describe how pure, dry barium sulfate can be obtained from the mixture. (2) (ii) Before certain X-rays, patients are given a suspension of barium sulfate to drink. Explain why this causes no harm to the patient, even though barium salts are toxic.	(iii) Explain how a sodium atom, Na, becomes a sodium ion, Na+.	(2)
	precipitate of barium sulfate forms in the mixture. The barium sulphate precipitate can be separated from the mixture and pur	ified.
	(Total for Question 2 =	· O marks)



Chemical reactions

3 (a) When ammonium nitrate dissolves in water, the change is endothermic.

In an experiment, a thermometer is placed in some water in a beaker and ammonium nitrate crystals are added.

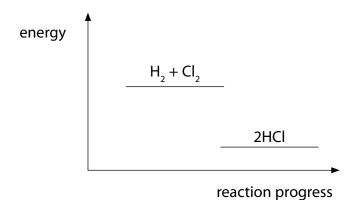
State what you would **see** to show that the reaction is endothermic.

(1)

(b) The reaction between hydrogen and chlorine to form hydrogen chloride is exothermic.

$$H_2 + Cl_2 \rightarrow 2HCl$$

The energy change taking place when the reactants, hydrogen and chlorine, form the product, hydrogen chloride, is shown in the diagram.



Explain, in terms of the breaking of bonds and the making of bonds, why this reaction is exothermic.

(2)

(c)	Marble chips react with dilute hydrochloric acid.	
	Marble is a form of calcium carbonate.	
	(i) Complete the balanced equation for this reaction.	(2)
	$CaCO_3$ +HCl \rightarrow $CaCl_2$ ++	
	(ii) Explain how using smaller sized marble chips affects the rate of this reaction, when all the other conditions remain the same.	(2)
	(iii) Explain, in terms of collisions between particles, how increasing the concentration of the hydrochloric acid affects the rate of this reaction, when all the other conditions remain the same.	
		(2)
	(Total for Question 3 = 9 ma	arks)



Atomic structure and isotopes

- **4** (a) An atom of potassium has an atomic number of 19 and a mass number of 39.
 - (i) Complete the table to show the number of protons, neutrons and electrons in this potassium atom.

(2)

	number of	
protons	neutrons	electrons

(ii)	Describe the	positions	of these	particles in	the	potassium atom.

(2)

(iii) State the electronic configuration of this potassium atom.

(1)

(iv) Give the relative mass of an electron.

(1)

Isotopes are different atoms of the same element that contain A different numbers of electrons B different numbers of protons but same number of electrons C same number of protons but different numbers of neutrons D same number of neutrons (ii) The sample of gallium contains 60.2% of gallium-69	
 ■ B different numbers of protons but same number of electrons ■ C same number of protons but different numbers of neutrons ■ D same number of neutrons (ii) The sample of gallium contains 60.2% of gallium-69 	
 C same number of protons but different numbers of neutrons D same number of neutrons (ii) The sample of gallium contains 60.2% of gallium-69 	
D same number of neutrons (ii) The sample of gallium contains 60.2% of gallium-69	
(ii) The sample of gallium contains 60.2% of gallium-69	
60.2% of gallium-69	
60.2% of gallium-69	
39.8% of gallium-71	
Calculate the relative atomic mass of gallium.	
(3)	
(Total for Question 4 = 10 marks)	





Chlorine

- **5** Chlorine is an element in group 7 of the periodic table.
 - (a) Chlorine, Cl₂, is a simple molecular, covalent substance.

The atoms in a molecule of chlorine are held together by a covalent bond.

(i) Explain what is meant by the term **covalent bond**.

(2)

(ii) Phosphorus reacts with chlorine to form phosphorus trichloride, PCl₃.

A phosphorus atom has five electrons in its outer shell.

A chlorine atom has seven electrons in its outer shell.

Draw the dot and cross diagram to show the bonding in a molecule of phosphorus trichloride, PCl₃.

Show outer electrons only.

(2)

(iii) Aluminium reacts with chlorine to form aluminium chloride, AlCl₃.

Write the balanced equation for this reaction.

(2)



*(b) Chlorine, Cl₂, is a simple molecular, covalent substance. Diamond is a giant molecular, covalent substance.

Sodium chloride is an ionic substance.

Zinc is metallic.

As a result of their different structures these substances have the following different properties.

- Solid chlorine has a very low melting point but diamond, sodium chloride and zinc have high melting points.
- Diamond and sodium chloride have different solubilities in water.

In terms of the structure and bonding of these substances, explain these properties.	
F F	(6)







(a) Calculate the percentage by mass of nitrogen in ammonium nitrate, N (relative atomic masses: $H = 1.0$, $N = 14$, $O = 16$)	, ,
	(3)
percentage by mass of nitrogen =	
(b) In the extraction of titanium from its ore, the final stage involves the between titanium(IV) chloride, TiCl ₄ , and sodium.	eaction
$TiCl_4 + 4Na \rightarrow Ti + 4NaCl$	
Calculate the maximum mass of titanium that can be obtained from 5 of titanium(IV) chloride in this reaction.	00 tonnes
(relative atomic mass: Ti = 48	
relative formula mass of $TiCl_4 = 190$)	(2)
mass of titanium =	tor
(c) Reactions used in the chemical industry often produce unwanted protogether with the required product.	ducts
State a problem this may cause for the manufacturer.	(1)



*(d) An experiment is carried out to determine the empirical formula of magnesium oxide.		
magnesium + oxygen → magnesium oxide		
The following results are obtained		
mass of magnesium ribbon reacted $= 0.420 g$ mass of magnesium oxide formed $= 0.700 g$		
Describe an experiment to produce these results. As part of your answer show how these results can be used to obtain the empirical formula of the magnesium oxide.		
(relative atomic masses: $Mg = 24.0$, $O = 16.0$)	(6)	
	(0)	

(Total for Question 6 = 12 marks)
·
TOTAL FOR PAPER = 60 MARKS



20